

PHY2822

Exp.5

Prelab ex

Balmer series of H

Electron transition: from  $n = 4$  to  $n = 2$ .

Wavelength  $\lambda = 486.1 \text{ nm}$

$$\text{Energy } E = h\nu = \frac{hc}{\lambda} = \frac{1240 \text{ eV} \cdot \text{nm}}{486.1 \text{ nm}} = 2.551 \text{ eV}$$